

# Polyol Modification of PEDOT: PSS as Hole Transport Material Affects the Performance and Stability of Calcium Titanate ( $\text{CaTiO}_3$ ) Solar Cell and UV Photodetector

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**Abstract.** This report a novel and straightforward architecture of  $\text{CaTiO}_3$ -based solar cells with optimized hole transport material (HTM) properties. The perovskite  $\text{CaTiO}_3$  particle was prepared via the solution process using different  $\text{CaCO}_3/\text{TiO}_2$  ratios. The design of  $\text{CaTiO}_3$ -based solar cell follows the inverted architecture, in which poly (4-styrene sulfonate)-doped poly (3,4-ethylenedioxy-thiophene (PEDOT:PSS) and PCBM-C70 were used as HTM and electron transport layer (ETL), respectively. Charge mobility of HTM was modified and improved by using either ethylene glycol (EG) or diethylene glycol (DEG). The optimum condition of  $\text{CaTiO}_3$ -based solar cell was obtained using  $\text{CaCO}_3/\text{TiO}_2$  ratio of (1:7) annealed at 900 °C and 44.4 % (v/v) DEG-modified PEDOT/PSS. The conversion efficiency of 3.1 % with a stable solar cell performance up to 72 h under the ambient condition without encapsulation was achieved.

**Keywords:** Electron transport layer, inverted structure, Perovskite  $\text{CaTiO}_3$  particle, perovskite solar cell, polyol alteration.

## 1 Introduction

Renewable energy utilization is considered an alternative solution to reduce environmental issues such as global warming and climate change due to the intensive use of fossil fuels. Among various renewable energy resources and conversion technology, solar energy

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conversion using photovoltaic (PV) has been a mature technology for future global green energy production [1, 2]. Thus, extensive researches have been geared to develop new materials for green energy applications. Perovskite solar cell (PSC) is an emerging solar cell technology that has already reached high efficiency since its seminal report [1]. Nonetheless, the current drawbacks of this solar cell are the use of lead halide perovskite, which features high environmental toxicity [1, 2]. Also, PSCs also show lower stability due to the degradation of organic molecules upon aging and irradiation. Therefore, one of the recent studies focuses on the development of Pb-free perovskite and stable solar cell performance.

Another direction to optimize the PSCs performance is to stabilize charge-separated states created upon light absorption by active perovskite layer: The excited state electron is donated to the electron transport layer (ETL), such as phenyl-C70-butyric acid methyl ester (PC70BM), while the hole is transferred to hole transport layer (HTL). Particularly for HTL poly(4-styrene sulfonate)-doped poly(3,4-ethylenedioxy-thiophene) (PEDOT:PSS) is intensively utilized in PSC [3, 4]. This material exhibits high conductivity, high transparency, and hole affinity [5–7]. Also, the molecular structure of PEDOT:PSS allows us to have a good photo and electrical-stability in an ambient atmosphere [7]. When integrated into PSC, a very high charge mobility is required for PEDOT:PSS since the hole needs to escape from recombination reaction with the electron residing on ETL and hence, creates more stable charge-separated states. Furthermore, the coating of PEDOT:PSS onto the photoactive layers remains to be one of the main technical challenges involved in the fabrication of these devices [8, 9].

In order to improve the properties in PEDOT:PSS, various chemical treatments have been carried out to control the electronic properties, which included the use of different organic solvents such as ethanol, dimethylsulfoxide, acetonitrile, and tetra-hydrofuran [10, 11]. Ouyang and co-workers have studied the effect of various salts on PEDOT:PSS conductivity, and the study revealed that the conductivity could be enhanced by employing metal salts as dopants due to the binding of the metal ions to the PSS anions [2, 12]. Alternatively, ethylene glycol (EG) have been explored as dopant material for PEDOT:PSS that is able to increase conductivity [13, 14]. Ethylene glycol and diethylene glycol (DEG) are both known as polyol material that plays a role in coordinating solvents, complexants, and surfactants which prevents agglomeration affecting electronic properties [15]. Nonetheless, the optimum doping condition of these EG and DEG modified PEDOT:PSS has not been reported.

Hence, this study aims at developing the novel architecture of inverted PSC and investigating the effect of modification on the hole transport materials, i.e., PEDOT:PSS, using EG and DEG at various doping concentration to the performance of the solar cell. The inverted PSC here was designed bearing  $\text{CaTiO}_3$  as the active layer since, among various perovskite materials,  $\text{CaTiO}_3$  constitutes an alternative to the lead-free based commercial perovskite because of its environmentally friendly properties [16, 17]. This material is well known for its high dielectric, luminescent, and semiconductor properties [18]. The optical properties of  $\text{CaTiO}_3$  can be tuned by modification of the precursors, which in this work a mixture of  $\text{CaCO}_3$  and  $\text{TiO}_2$  at different ratios was used for preparation *via* wet chemical routes. Finally, solar cell performance by combining modified PEDOT:PSS and  $\text{CaTiO}_3$  as well as photoresponse as UV photodetector is presented.

## 2 Experimentals

### 2.1 Materials and synthesis of $\text{CaTiO}_3$

The materials used were the commercial  $\text{CaCO}_3$ , and anatase  $\text{TiO}_2$  powder (MTI, 99 %), Solvents used were 96 % ethanol (analytical grade) and distilled water. Mixtures of  $\text{CaCO}_3$  and  $\text{TiO}_2$  powder with  $\text{CaCO}_3/\text{TiO}_2$  molar ratio of 1:1, 1:3, and 1:7 were prepared. A mixture was dissolved in 100 mL of ethanol and homogenized by continuous stirring at 300 rpm (rpm = 1/60 Hz) for 2 h at room temperature. The solution was then washed with distilled water several times and dried in an oven at 100 °C for 2 h to obtain the white powder. The resulting white powder was subsequently ground into fine and homogeneous granules. Finally, the powder was annealed at 700 °C and 900 °C for 2 h.

### 2.2 Fabrication of $\text{CaTiO}_3$ solar cell and UV photodetector

Indium-doped tin oxide (ITO) on glass substrate was used as active layer to prepare thin film. For electron transport layer (ETL), PCBM-C70 was dissolved into chlorobenzene, dropped using micropipet onto ITO glass, and spin coated at 5 000 rpm for 30 s. The ETL layer was then dried on the heating plate at 120 °C for 15 min. As prepared  $\text{CaTiO}_3$  was dissolved into ethanol (EtOH) to form paste/slurry. The  $\text{CaTiO}_3$  was dropped and spin coated at 5 500 rpm for 30 s and dried on the heating plate at 120 °C for 10 min. The obtained layers were then coated with either pristine PEDOT:PSS or modified PEDOT:PSS at 7 000 rpm for 30 s. Modification of PEDOT:PSS was carried out by using either ethylene glycol (EG) or diethylene glycol (DEG). As cathode, carbon-coated ITO glass was used.

### 2.3 Characterization

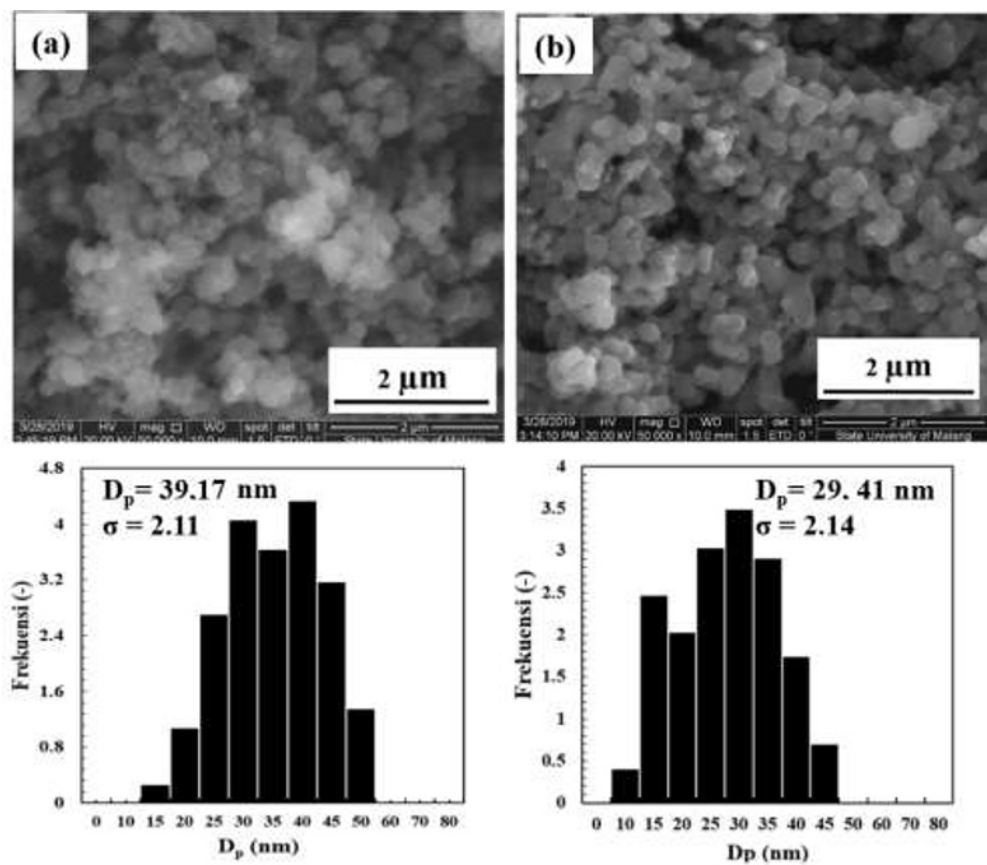
X-ray diffraction (XRD) pattern was obtained using a diffractometer (PAN analytical type X'Pert Pro) operated at 40 kV, and 40 mA with  $\text{Cu-K}\alpha$  as a radiation source. Samples were scanned between 10° and 100° (2 $\theta$ ) with resolutions of 0.05°. Micromorphology of  $\text{CaTiO}_3$  powder was observed by scanning electron microscopy (SEM, FEI Inspect 21 type) at 100 kV accelerating voltage. Electrochemical impedance characterization of modified PEDOT:PSS was undertaken using potentiostat (VerstaStat II) in two electrode system. FTIR spectra were measured using Thermo-Nicolet FTIR spectrometer. Absorption properties were measured using UV/vis spectrometer (Rayleigh UV-9200). The photovoltaic performance of the  $\text{CaTiO}_3$  solar cells was determined by measuring photocurrent density-photovoltage ( $J$ - $V$ ) characteristic curve under diffuse UV-vis irradiation. This  $J$ - $V$  measurement was implemented in ambient atmosphere by using an electrochemical measurement system to control the voltage and current. The light-to-electrical energy conversion efficiency ( $\eta$ ) of the solar cell were calculated according to the Equations (1):

$$\eta = \frac{FF \times V_{oc} \times I_{sc}}{P_{in}} \quad (1)$$

where  $I_{sc}$  is the short-circuit current (mA),  $V_{oc}$  is the open-circuit voltage (V),  $P_{in}$  is the incident light power (Watt), and FF is the fill factor.

3 Results and discussions

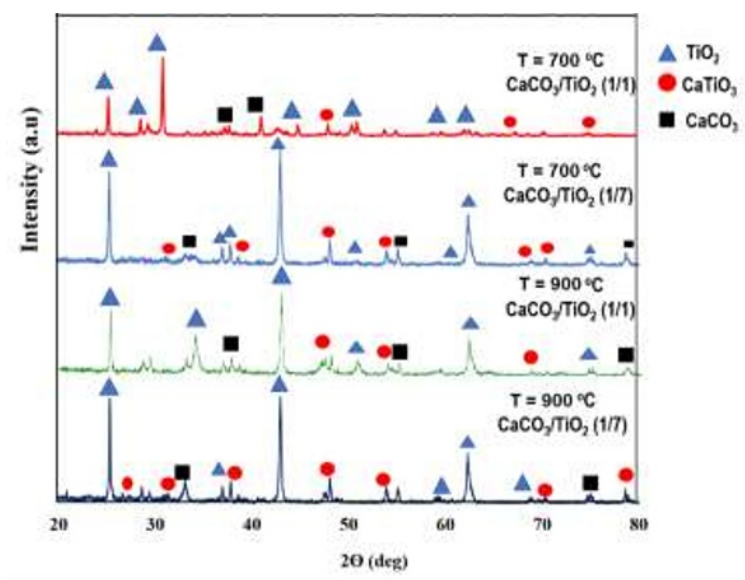
Figure 1 shows SEM images of CaTiO<sub>3</sub> with composition variation of CaCO<sub>3</sub>/TiO<sub>2</sub> (1:1 and 1:7) upon calcination at 900 °C. It is clear that the morphology of CaTiO<sub>3</sub> particle is semi-spherical with non uniform particle size and is relatively aggregated. The particle size distribution indicates that increasing molar composition of TiO<sub>2</sub> in CaCO<sub>3</sub>/TiO<sub>2</sub> from (1:1) to (1:7) under high calcination temperature yields smaller crystallite size. In addition, the resulting CaTiO<sub>3</sub> particles prepared using higher TiO<sub>2</sub> content tend to be more stable than that using lower TiO<sub>2</sub> content as well as yield less aggregated morphology. In contrast with the CaTiO<sub>3</sub> prepared *via* solid-state reaction, the heat treatment at high-temperature capable of changing the structural properties of wet-chemically synthesized CaTiO<sub>3</sub> as indicated by inhomogeneity in its final forms [19].



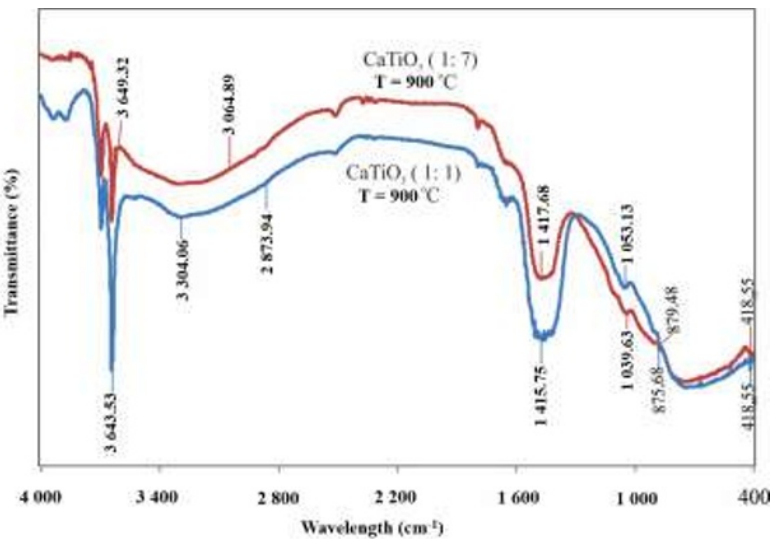
**Fig. 1.** Morphology of CaTiO<sub>3</sub> with different compositions of CaCO<sub>3</sub>/TiO<sub>2</sub> (a) 1:1, T = 900 °C and (b) 1:7, T = 900 °C.

Electronic properties of CaTiO<sub>3</sub> is further assessed from the X-ray diffraction (XRD) pattern. Figure 2 shows the resultant XRD spectra of CaTiO<sub>3</sub> prepared under different conditions, which are referenced to the standard peak intensities of CaCO<sub>3</sub>, TiO<sub>2</sub>, and CaTiO<sub>3</sub>. The decrease of peak intensities (diffraction maximum) as well as the peak broadening in CaCO<sub>3</sub> indicates the deformation of its crystal structure, i.e., crystal structure defects induce high crystal lattice strain [20]. These broadening peaks also indicate that

CaCO<sub>3</sub> and TiO<sub>2</sub> have lower crystallinity degrees. After calcination, X-ray diffraction pattern shows that the CaTiO<sub>3</sub> is obtained in all various samples as new peaks at 33.17°, 41.32°, 48.11°, and 75.11° appear which can be referenced to JCPDS No. 96-101-1212 for CaTiO<sub>3</sub>. The results show that CaTiO<sub>3</sub> powder is characterized with orthorhombic structure. In addition, as shown in Figure 2, there exist diffraction peaks of TiO<sub>2</sub> and CaCO<sub>3</sub> in the CaTiO<sub>3</sub> samples indicating that the preparation route was either excessive of precursor concentration or insufficient for mechanical grinding to provide fine precursor nanoparticle.



**Fig. 2.** XRD patterns of CaTiO<sub>3</sub> powders with different heating treatment temperature and compositions of CaCO<sub>3</sub> and TiO<sub>2</sub> for the comparison.



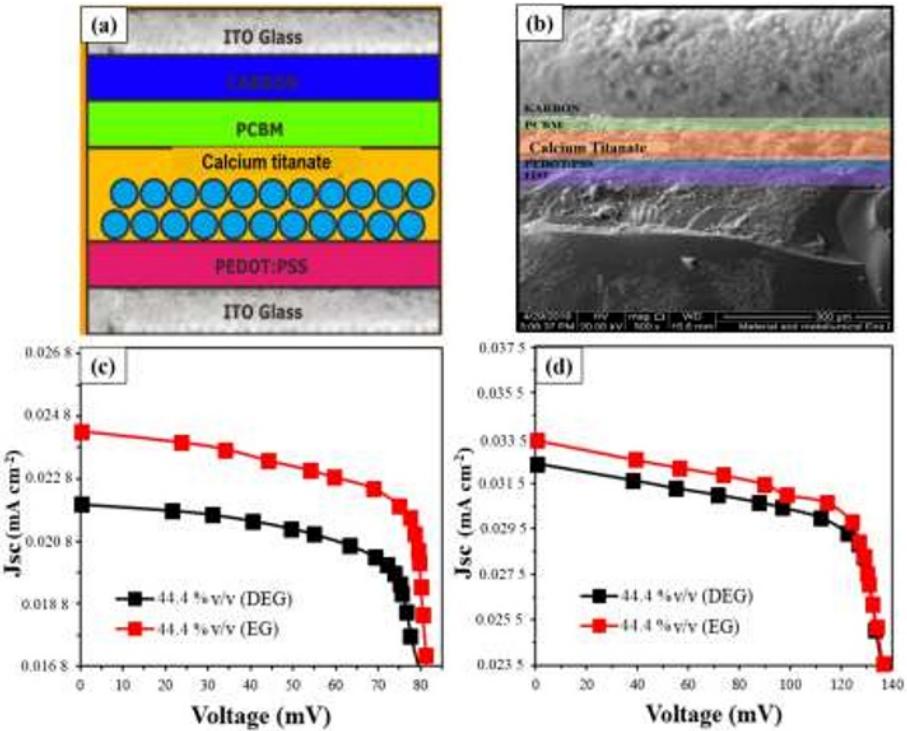
**Fig. 3.** FTIR spectra of CaTiO<sub>3</sub> prepared using different ratio of CaCO<sub>3</sub>/TiO<sub>2</sub>.

In addition, the functional group of CaTiO<sub>3</sub> were observed by FTIR analysis, and the results were shown in Figure 3. The IR spectra of CaTiO<sub>3</sub> sample with a varied

CaCO<sub>3</sub>/TiO<sub>2</sub> ratio (1:1) and (1:7) annealed at 900 °C show bands at 3 698 cm<sup>-1</sup> and 3 642 cm<sup>-1</sup>. According to the previous report, these peaks are assigned to a vibration characteristic of the hydroxy (OH) group. Moreover, it is indicated that increasing annealing temperature affects the absence of water molecules. The higher calcination temperature (T = 900 °C) can reduce the moisture content of material. The IR spectra also exhibit band at 1 422 cm<sup>-1</sup> which decreases with increasing TiO<sub>2</sub> fraction. According to the previous report, the IR band at 1 400 cm<sup>-1</sup> shows the asymmetrical and symmetrical vibrations between metal oxide. Hence, it indicates the residual interaction of CaCO<sub>3</sub> functional group, which is the bond vibration between C-O of CO<sub>3</sub><sup>2-</sup> ions. The increasing mol ratio of TiO<sub>2</sub> increases the absorption area leading to more surface reaction between TiO<sub>2</sub> and CaCO<sub>3</sub>. Therefore, the vibrational spectrum at 1 066 cm<sup>-1</sup> shows the bond vibration between C-O-Ti groups whilst the IR peaks around 800 cm<sup>-1</sup> to 900 cm<sup>-1</sup> shows the existence of Ti-O-Ti bond vibration. In addition, the broad absorption at 664 cm<sup>-1</sup> indicates the vibration of the Ca-Ti-O group.

**Table 1.** Estimated of optical band gap for CaTiO<sub>3</sub> prepared using different CaCO<sub>3</sub>:TiO<sub>2</sub> ratio and annealing temperature.

CaCO <sub>3</sub> :TiO <sub>2</sub> ratio	Eg/eV (T = 700 °C)	Eg/eV (T = 900 °C)
1:1	3.98	3.69
1:3	4.36	3.95
1:7	4.29	4.22



**Fig. 4.** (a) The device architecture of solar cell, (b) cross-sectional SEM image of the CaTiO<sub>3</sub> solar cell. The current density-voltage characteristic (*J*-*V*) curve of CaTiO<sub>3</sub> solar cell bearing modified EDOT:PSS and different composition of CaCO<sub>3</sub>/TiO<sub>2</sub> (c) 1:1 and (d) 1:7 annealed at 900 °C.

As discussed earlier, increasing annealing temperature as well as concentration of TiO<sub>2</sub> will cause the growth of CaTiO<sub>3</sub> crystal. This altered crystallite CaTiO<sub>3</sub> structure as

revealed in XRD analysis in turn affects its optical properties. The optical properties of  $\text{CaTiO}_3$  are indicated by optical band gap estimated from absorption spectra (using Tauc Plot, Table 1). As seen, it is obvious that  $\text{CaCO}_3/\text{TiO}_2$  ratio and annealing temperature affect the optical band gap, in which higher  $\text{TiO}_2$  ratio yields higher optical band gap. Meanwhile, higher annealing temperature results in a slightly lower optical band gap due to impaired quantum confinement effects. From this point, we will explore the utilization of  $\text{CaTiO}_3$  prepared at 900 °C using  $\text{CaCO}_3/\text{TiO}_2$  ratio (1:1) and (1:7) because of their significant different in optical properties.

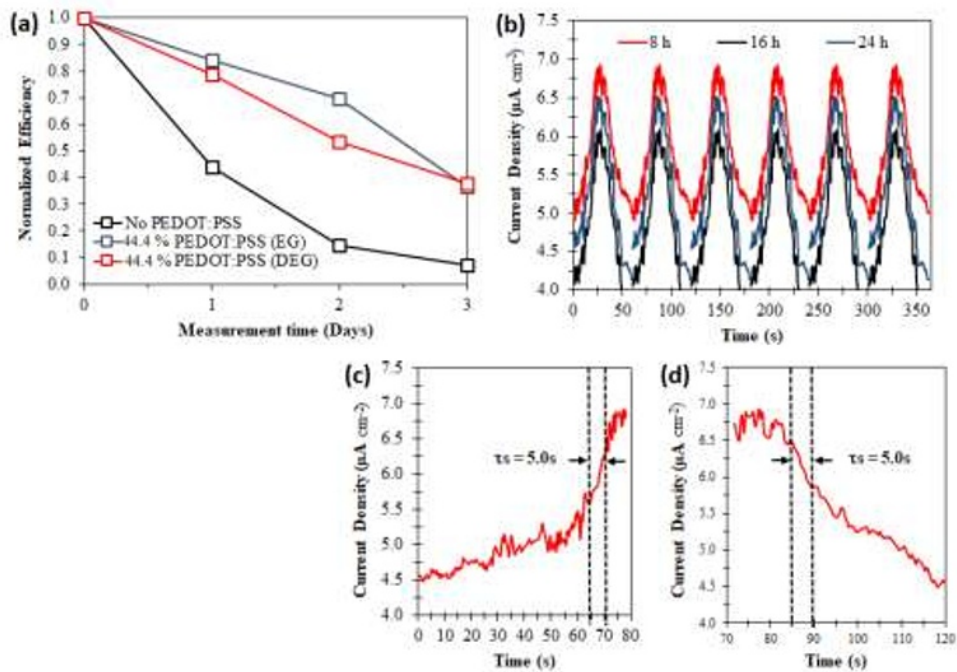
**Table 2.** Solar cell properties derived from current-voltage measurements.

$\text{CaCO}_3:\text{TiO}_2$	Modification	Jsc (mA cm <sup>-2</sup> )	Voc (mV)	FF	η (%)
1:1	EG – 16.7 %	0.017	60	0.66	0.64
1:1	EG – 28.6 %	0.019	68	0.64	0.76
1:1	EG – 37.5 %	0.020	71	0.66	0.92
1:1	EG – 44.4 %	0.025	81	0.68	1.32
1:1	DEG – 16.7 %	0.017	62	0.65	0.65
1:1	DEG – 28.6 %	0.020	72	0.65	0.92
1:1	DEG – 37.5 %	0.020	81	0.69	1.08
1:1	DEG – 44.4 %	0.025	84	0.69	1.42
1:7	EG – 16.7 %	0.024	90	0.75	1.54
1:7	EG – 28.6 %	0.027	119	0.60	1.82
1:7	EG – 37.5 %	0.027	119	0.66	2.02
1:7	EG – 44.4 %	0.032	136	0.67	2.58
1:7	DEG – 16.7s %	0.025	82.2	0.79	1.57
1:7	DEG – 28.6 %	0.027	99.2	0.73	1.90
1:7	DEG – 37.5 %	0.032	105	0.64	2.04
1:7	DEG – 44.4 %	0.032	136	0.67	3.11

Figure 4a and Figure 4b show the architecture design and the cross-sectional image of  $\text{CaTiO}_3$  solar cell, respectively. The solar cell architecture reflects an inverted structure of perovskite solar cell, in which PEDOT:PSS and PCBM-C70 are employed as hole and electron transport material, respectively. Performance of  $\text{CaTiO}_3$  perovskite solar cells, including cell parameters (open circuit voltage Voc, short circuit current density Jsc, and fill factor FF), is evaluated from current density–voltage (J-V) curves measured under diffuse irradiation (~380 W m<sup>-2</sup>, incandescent lamp). As shown in Figure 4c and Figure 4d as well as in Table 2, polyol modification does not significantly alter Voc. However, substitution of  $\text{CaTiO}_3$  (1:1) to (1:7) enhances Voc by 50 % to 100 %, which is due to a higher optical band gap for  $\text{CaTiO}_3$  (1:7) than  $\text{CaTiO}_3$  (1:1). Interestingly, the deployment of EG-doped PEDOT:PSS in  $\text{CaTiO}_3$  solar cells consistently shows slightly lower Jsc than that of DEG-doped PEDOT:PSS. The higher Jsc obtained in solar cell using DEG-doped PEDOT:PSS can be due to the increased hole mobility in PEDOT:PSS: The impedance characterization reveals that hole mobility in PEDOT:PSS (0.45 cm<sup>2</sup> V<sup>-1</sup> s<sup>-1</sup>) is increased by factor of 1.6 and 2.3 upon modification using EG and DEG, respectively. This results is consistent with literature report [13, 14] that EG doping in PEDOT:PSS increases the electron phonon coupling and hence its conductivity. It is known that higher mobility of HTM should lead to a suppressed recombination reaction, a limiting process to generate

high photocurrent in solar cell. Overall, the best photovoltaic performance in this work exhibit a conversion efficiency ( $\eta$ ) as high as 3.11 %.

As central issue of perovskite solar cell is its stability, the effect of PEDOT:PSS modification toward stability of  $\text{CaTiO}_3$  solar cell efficiency is assessed (Figure 5a) upon aging for 3 d under ambient condition with 60 % relative humidity. Without polyol modification, solar cell efficiency drops down to 7 % of its initial magnitude. Polyol modification of PEDOT:PSS, irrespective of dopant materials, enables maintaining 40 % of initial efficiency. Nevertheless, it should be noted that solar cells using EG-doped PEDOT:PSS reveal a slower degradation rate. To further confirm the origin of photo response, the optoelectrical properties have been investigated as UV photodetector (Figure 5b). The present device exhibits a sharp on/off transition in the photocurrent with good stability and reproducibility upon UV-lamp exposure (365 nm,  $\sim 100 \text{ mW cm}^{-2}$ ). The photocurrent different between dark and UV irradiated is  $2.5 \text{ }\mu\text{A cm}^{-2}$ , i.e., one order of magnitude lower than the obtained  $J_{sc}$ . This result is plausible since the perovskite  $\text{CaTiO}_3$  strongly absorbs deep UV region  $< 300 \text{ nm}$ . In addition, the transient behaviors of current density of  $\text{CaTiO}_3$  solar cell (Figure 5c and Figure 5d) is investigated to indicate the sensitivity. The rising and decaying time is found *ca.* 5 s, which is relatively slow for rapid UV photodetection. Therefore, further optimization is required, particularly to enhance the sensitivity of photodetection.



**Fig. 5.** (a) Efficiency of  $\text{CaTiO}_3$  solar cells and (b) chopped light UV photocurrent response upon aging for days and hours, respectively. Rising (c) and decaying (d) current response upon UV irradiation.

## 4 Conclusion

This research have demonstrated that the novel inverted architecture of  $\text{CaTiO}_3$ -based solar cell using PEDOT:PSS as hole transport material functionally works. Polyol modification by either EG or DEG doping significantly enhances the charge mobility in PEDOT:PSS and affects the current generation in solar cell upon irradiation. The optimum conditions of  $\text{CaTiO}_3$  solar cell was achieved by using  $\text{CaTiO}_3$  particles prepared using  $\text{CaCO}_3/\text{TiO}_2$  ratio of (1:7) annealed at 900 °C and 44.4 % DEG-modified PEDOT:PSS. The power efficiency obtained under diffuse illumination is 2.58 %. The improvement in the electrophysical properties of PEDOT:PSS as the hole transport layer combined with PCBM-C70 proved to be key factors towards enhancing the  $\text{CaTiO}_3$  solar cells. These improvements include a significantly improved stability of power conversion efficiency upon aging 72 h. In addition, the novel design also performs sensitive UV photoresponse even upon irradiation at UV region with lower energy excitation.

Technical assistance from Central Mineral and Advanced Material Laboratory of Malang State University and Bandung Institute of Technology for samples characterization is gratefully acknowledged. This work was financially supported by Science Technology Ressearch Grant, Indonesia Toray Science and Foundation (ITSF).

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